Original research

Synthesis, Characterization and Antibacterial Assay of Some Schiff Base Metal (II) complexes

Abstract

The complexes of Co (II), Mn (II) and Ni (II) with schiff base derived from pentane-2,4-dione and 2-aminobenzoic acid were synthesized and characterized by molar conductivity, magnetic susceptibility, Infrared and elemental analyses. The solubility test on the schiff base and its metal (II) complexes revealed their solubility in most organic solvents except chloroform and diethyl ether. The molar conductivity of the complexes was high indicating that they are strong electrolytes. The antibacterial susceptibility test conducted on the schiff base and the metal (II) complexes showed a good activity except Ni (II) complex.

Keywords: Antibacterial susceptibility, electrolytes, magnetic susceptibility, molar conductivity, Schiff base ligands.

Introduction

The name coordination compound comes from the coordinate covalent bond, which historically was considered to form by donation of a pair of electrons from one atom to another. In coordination compounds the donors are usually the ligands, and the acceptors are the metals. Coordinate covalent bonds formally formed by combining one electron from each atom; only the formal electron counting distinguishes them [1]. A coordination compound consists of a central atom or ion, which is usually metallic and is called the coordination centre, and a surrounding array of bound molecules or ions, that are in turn known as ligands or complexing agent [2].

Schiff bases are condensation products of primary amines with carbonyl compounds. They were first reported by Schiff in 1864, [3]. The common structural feature of these compounds is the azomethine group with the general formula RHC=NR', where R' and R is alkyl, aryl, cyclo alkyl or heterocyclic groups which may be variously substituted [4]. As a result of the relative simple preparation, synthetic flexibility, and the special property of azomethine group(C=N), Schiff bases are generally excellent chelating agents [4], especially when a functional group like -OH or –SH is present in the position close to the azomethine group so as to form five or six membered ring with a metal ion. Versatility of Schiff base ligands and biological, analytical, and industrial applications of their complexes make further investigations in this area highly desirable [4]. Thus, the research focused on the synthesis, characterization and exploring antimicrobial activity of the Schiff base and the complexes of cobalt (II), nickel (II), and manganese (II).

Materials and Method

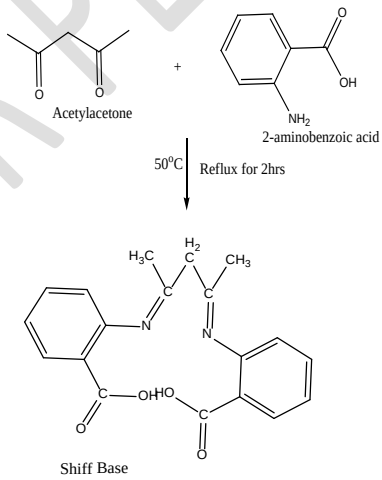
Materials

All chemical reagents, 2-Aminobenzoic acid, acetylacetone, and solvents were of analytical grade were obtained from LOBA Chemie, Park Scientific Ltd, UK and JHD and used without any purification. The microorganisms (clinical isolates) used for the antimicrobial analysis were obtained from Microbiology Department, Usmanu Danfodiyo University, Sokoto.

Synthesis of the Schiff base and the metal complexes

The schiff base was synthesized by the of addition of 25cm^3 of (1.0269 cm^3, 0.01 mol) acetyl acetone to the same volume of (2.7428 g, 0.02 mol) 2-aminobenzoic acid ethanolic solution. The resultant mixture was refluxed for two
at about 50 °C. The solution was concentrated on steam bath and then allowed to cool in an ice bath. The dark brown product that precipitated was recrystallized from hot ethanol. The crystals were then filtered and dried in a desiccator over phosphorous pentoxide [5].

The cobalt complex was prepared by adding 25 cm^3 of ethanolic solution of cobalt chloride (2.3393 g, 0.01 mol) with ethanolic solutions of the prepared Schiff base (3.3864 g, 0.01 mole). The resulting mixture was refluxed for two hours after which the solution was concentrated on a steam bath and cooled in ice cold water. The precipitate was filtered and washed repeatedly with hot ethanol until the washing was colourless. The obtained product was dried in a desiccator over phosphorus pentoxide [5,6]. The same procedure was adopted for the synthesis of manganese and nickel complexes.



Scheme 1: Proposed structure of the synthesized Schiff base of acetylaceto-2-aminophenoic acid

Characterization of the Schiff base and the complexes

Fourier-transform infrared spectroscopy (FTIR) and elemental analyses

The FTIR spectra of the **Schiff base** and complexes were recorded in the range of 4000 - 650cm^{-1} using Cary 630 FTIR spectrometer at the Department of Chemistry, Bayero University, Kano. The percentage mass of carbon, hydrogen, nitrogen and oxygen were determined by using PerkinElmer CHNS elemental analyser at the Universiti Tecknologi Petronas (UTP), Malaysia.

Determination of melting point, molar conductivity, magnetic susceptibility and solubility test

The melting points of the ligand and the complexes are uncorrected as determined by Gallenkemp melting point apparatus. For the molar conductivity, a solution of each metal (II) complex (0.02 g/ml) was prepared in dimethyl sulfoxide and the molar conductance was measured using the Jenway conductivity meter. The molar conductance of the complexes was obtained using the relation:

$$\text{Molar Conductance} = \frac{1000K}{C} \quad \ldots\ldots\ldots\ldots\ldots\ldots\ldots\ldots\ldots\ldots\ldots\ldots\ldots\ldots 1$$

where K = specific conductance, C = molar concentration.

The magnetic susceptibility of the complexes is determined by placing an empty capillary tube (W_o) inside the magnetic susceptibility balance, and the value recorded as R_o. Small quantity of the complex was then placed in to the capillary tube (W_1) and, the value R and length L of the sample in the tube are also recorded respectively. The magnetic susceptibility (Xg) was obtained using the relation:

$$Xg = \frac{CL(R-R_o)}{10^9 M} \quad \ldots\ldots\ldots\ldots\ldots\ldots\ldots\ldots\ldots\ldots\ldots\ldots\ldots\ldots 2$$

Where M is the mass of complex in the capillary tube (W_1-W_o) and C is the proportionality constant which is always 1.

The solubility of the **Schiff base** and the metal complexes were carried out in distilled water, methanol, dimethyl sulfoxide, hexane, ethanol, acetone, and diethyl ether. 10 mg of each of the metal complex was dissolved in 2 mL of corresponding solvent to determine their solubility [7].

Antibacterial studies of the Schiff base and metal (II) complexes

The *in vitro* antibacterial were tested against two pathogenic bacteria: *Streptococcus pyogenes* (gram positive) *and Pseudomonas aeruginosa* (gram negative). The bacteria were sub cultured on nutrient agar media and incubated at 37 °C for 24 hours. Three concentrations (50, 100, and 150 mg/mL) of both the **Schiff base** and the complexes were made using DMSO solvent. Ditch Well diffusion method was used to assay the antibacterial activity.

Preparation of media for antibacterial test

A 2.8 g of nutrient agar was weighed and dissolved in 100 mL distilled water in a conical flask which was then heated on a hot plate to dissolve the powder completely followed by autoclaving at 121 °C for 15 minutes. It was then allowed to cool and poured into Petri dishes to solidify.

Preparation of MacFarland turbidity standard solution

The turbidity standard was prepared by mixing 99.5 cm^3 of 1% v/v sulphuric acid and 0.5 cm^3 of 5 mM barium chloride ($BaCl_2.2H_2O$). The solution was mixed thoroughly and dispensed into the test tubes [8].

Antibacterial assay

The microorganisms (clinical isolates) used for the antimicrobial analysis were obtained from Microbiology Department, Usmanu Danfodiyo University, Sokoto.

The antibacterial effect of the complexes and the Schiff base was performed using the procedure described by [9]. A sterile Muller Hinton agar was prepared and poured into the Petri dishes and allowed to solidify. A loop full of each organism was stricken on the surface of the solidified media and each plate properly labeled. Wells are made in each plate using a sterile cork borer (6 mm), after which the concentration of each extract was dispensed into the bored holes alongside the antibiotic (streptomycin) which served as the control. The plates were left to stand to allow diffusion of the extract after which it was incubated at 37 °C for 24 hours. The diameter of zone of inhibition was measured and recorded using a meter rule. These activities were performed three times and reported as mean of all the three readings.

Results and Discussion

The results of the various analyses carried out are presented below. The interaction between acetyl acetone and 2-aminobenzoic acid gives a shiny brown Schiff base of 68% yield with a melting point of 213 °C, indicating good stability. The melting point might be associated with the strong attractive forces due to intermolecular hydrogen bond. The reaction mixture of the Schiff base and the metal chloride of Mn (II), Co (II), and Ni (II) in ethanol yielded 61.53 - 73.13% of the metal (II) complexes. They were isolated as crystals stable in atmospheric condition and are pale brown, brown, and green colours, with a melting point ranging between 97 to 165 °C as shown in Table 1 [10].

Table 1: Some physical properties of the ligand and the complexes

Complexes	Colour	Yield (%)	M.W	M.P/ D.T (°C)
L^1	Brown	68.00	338.36	184

[Mn(L^1)].4H_2O	Pale Brown	61.53	465.32	165
[Co(L^1)].2H_2O	Green	88.57	433.33	97
[Ni(L^1)].4H_2O	Pale Brown	73.13	469.07	107

M.W = Molecular weight, M.P/ D.T = Melting point/ Decomposition temperature.

The band at 1615 cm^{-1} in the Schiff base spectral data was as assigned to stretching vibration mode of *v*(C=N). The spectral of the metal (II) complexes assignable to *v*(C=N) vibration mode, undergoes a shift to lower wave number in the range of 1588 - 1598 cm^{-1} on coordination. The band within 409 - 431 cm^{-1} and 455 – 480 cm^{-1} are attributed to *v*(M-N) and *v*(M-O) stretching vibrations respectively, confirming coordination of the **Schiff base** to the respective metal ions (Table 2). The bands in the region 3371 - 3473 cm^{-1} was attributed to *v*(O-H) stretching frequency for water of crystallization in the metal (II) complexes [11,12]. The elemental analyses (C, N, and H) of the **Schiff base** and the complexes were determined and presented in Table 3. The result obtained is in good agreement with the 1:1 metal to **Schiff base** ligand ratio.

Table 2: Infrared spectral data of the ligand and the complexes

Compounds	*v* (C=N)	*v* (C=O) assy.	*v* (C=O) symm	*v* (C=C)	*v*(O-H)	*v*(M-N)	*v*(M-O)
L^1							
[Mn(L^1)].4H_2O	1588	1538	1390	1460	3371	480	431
[Co(L^1)].2H_2O	1592	1544	1396	1469	3393	455	422
[Ni(L^1)].4H_2O	1598	1540	1394	1467	3473	474	409

Table 3: Elemental results of the complexes and the ligands (%).

	Calculated (found)		
Complexes	**C**	**H**	**N**
L^1	67.45(67.40)	5.36(5.40)	8.28(8.22)
[MnL].4H_2O	49.04(49.10)	5.63(5.70)	6.02(6.08)
[CoL].2H_2O	52.67(52.60)	5.12(5.20)	6.46(6.50)
[NiL].4H_2O	48.65(48.69)	5.59(5.62)	5.97(6.03)

Table 4 showed magnetic susceptibility values for the metal (II) complexes in the range of $6.1x10^{-6}$ to $2.5x10^{-5}$g. They are all positive values, suggesting that the complexes are paramagnetic in nature. The molar conductance values of the metal (II) **Schiff base** complexes in DSMO solution were determined in the range of 7.5 – 14.6 $\Omega^{-1}cm^{-2}mol^{-1}$ (Table 5). These values are small thus, suggesting that the metal (II) **Schiff base** complexes are non-electrolyte [13].

Table 4: Magnetic properties of the metal (II) complexes

Complex	R_o	R	L/cm	M (W_1-W_o) / g	Xg/cm
$[MnL].4H_2O$	-037	946	1.900	0.075	$2.5x10^{-5}$
$[CoL].2H_2O$	-038	183	2.000	0.086	$1.6x10^{-5}$
$[NiL].4H_2O$	-037	720	1.800	0.073	$6.1x10^{-6}$

W_1 = Weight of capillary tube with metal complex inside the magnetic susceptibility balance and W_o = weight of empty capillary tube inside the magnetic susceptibility balance.

Table 5: Conductivity measurement of the metal (II) complex in DMSO

Complexes	Concentration (M)	Molar conductivity ($\Omega^{-1}cm^{-2}mol^{-1}$)
$[Mn(L^1)].4H_2O$	$2x10^{-2}$	14.6
$[Co(L^1)].2H_2O$	$2x10^{-2}$	13.5
$[Ni(L^1)].4H_2O]$	$2x10^{-2}$	7.5

The solubility test of the Schiff base and its metal (II) complexes were carried out in various solvents. The Schiff base is soluble in acetone, methanol, DMSO, ethanol, and hexane, slightly soluble in diethyl ether but insoluble in water and chloroform. The metal (II) complexes are soluble in acetone, DMSO, ethanol and hexane but slightly soluble in chloroform and diethyl ether. Mn (II) and Co (II) complexes were slightly soluble in distilled water, chloroform and diethyl ether, whereas Ni (II) complex is insoluble in water but slightly soluble in diethyl ether as shown in Table 6.

Table 6: Solubility test of the Schiff base and the complexes

Solvent	Ligand	$[MnL].4H_2O$	$[CoL].2H_2O$	$[NiL].4H_2O$
Water	NS	SS	SS	NS
Acetone	S	S	S	S
Methanol	S	S	SS	S
Chloroform	NS	SS	SS	SS
DMSO	S	S	S	S
Ethanol	S	S	S	S
Hexane	S	S	S	S
Diethylether	SS	SS	SS	SS

S = Soluble, NS =Not Soluble, SS=Slightly Soluble.

The antibacterial test carried out showed a good activity on the ligand and two of the complexes of Co (II) and Mn (II)] (Figures 1 and 2). However, the Ni (II) complex did not show any activity on the bacteria used.

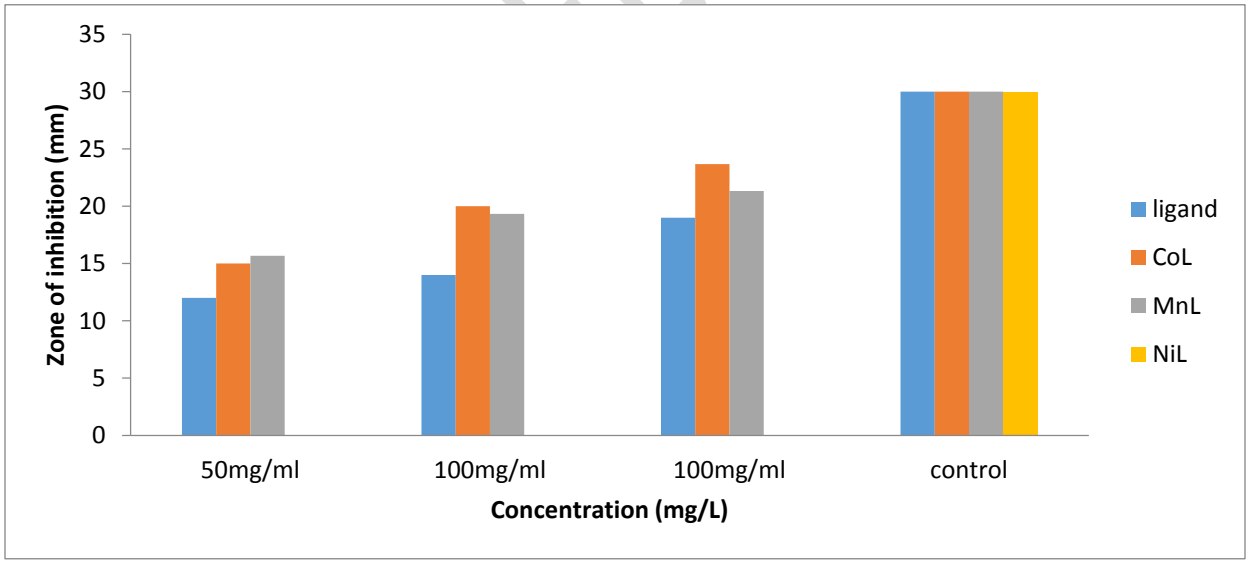


Fig. 1: Antibacterial activities of the Schiff base and complexes against clinical isolated *Streptococcus Pyogens.*

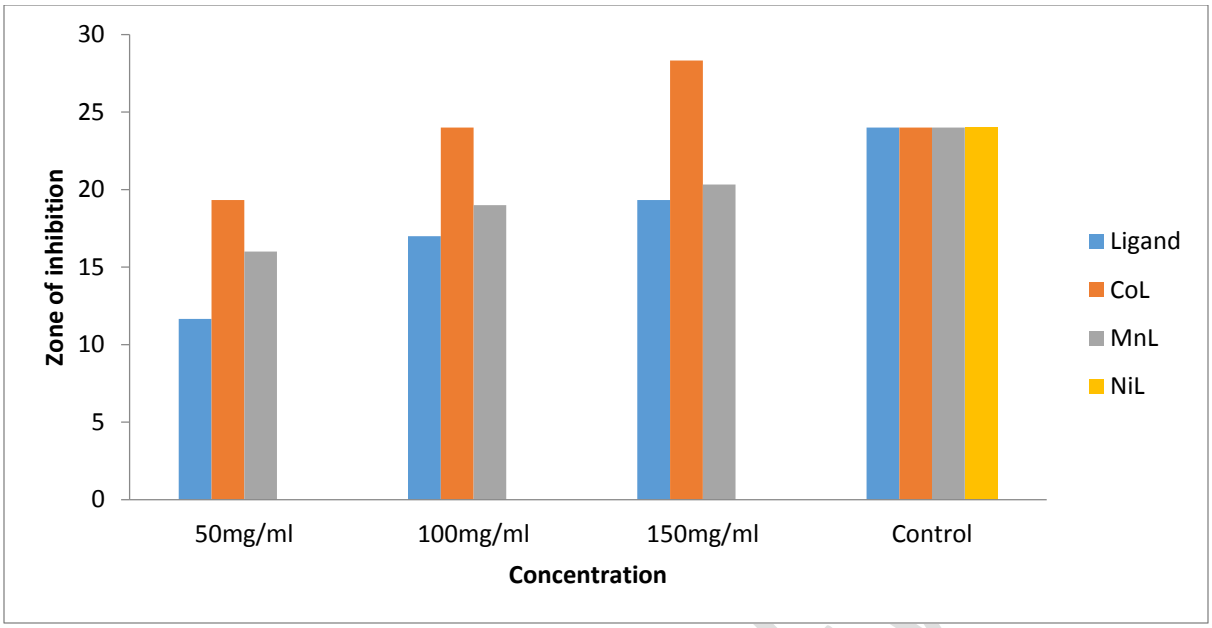


Fig. 2: Antibacterial activities of the Schiff base and complexes against clinical isolate *P. aeruginosa*.

Conclusion

The Schiff base (L^1) of 2-aminobenzoic acid and acetyl acetone, and its Mn (II), Co (II) and Ni (II) were successfully synthesized and characterized. The conductivity measurement of the complexes showed a very high value revealing that they are very good electrolytes. The solubility test was carried out in different solvents and coordination in the complexes occurs through the N atom of the amine and also through the deprotonated O atom of the OH group from -COOH. The IR spectral data of the Schiff base confirm the existence of C=N group, suggesting the formation of the Schiff base. The synthesized Schiff base and the metal complexes displayed a good broad-spectrum antimicrobial activity against gram-positive and gram-negative bacteria at different concentrations except Ni (II) complex. Thus, L^1, $[Mn(L^1)].4H_2O$ and $[Co(L^1)].2H_2O$ should be considered as possible lead compounds to be developed into antibiotics against the tested bacterial strains *S. pyrogens*. *P. aeruginosa.*

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